



$$C \quad 12.0 \times 14 = 168.0$$

$$H \quad 1.0 \times 20 = 20.0$$

$$N \quad 14.0 \times 2 = 28.0$$

$$S \quad 32.1 \times 1 = 32.1$$

$$O \quad 16.0 \times 4 = 64.0$$

$$312.1 \text{ g/mol}$$

$$C \quad \frac{168.0 \text{ g}}{312.1 \text{ g}} \times 100 = 53.8\% C$$

$$H \quad \frac{20.0 \text{ g}}{312.1 \text{ g}} \times 100 = 6.4\% H$$

$$N \quad \frac{28.0 \text{ g}}{312.1 \text{ g}} \times 100 = 9.0\% N$$

$$S \quad \frac{32.1 \text{ g}}{312.1 \text{ g}} \times 100 = 10.3\% S$$

$$O \quad \frac{64.0 \text{ g}}{312.1 \text{ g}} \times 100 = 20.5\% O$$

add to 100.0% ✓
Check

